

Chapter 8 Covalent Bonding Study Guide Answers

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CHAPTER FOR Class Section 8.2 continued ... Differentiate between an ordinary covalent bond and a coordinate covalent bond. Give an example of a molecule that exhibits both and label them.

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Chapter 8 Concepts of Chemical Bonding . Chemical Bonds Three types: - Ionic Electrostatic attraction between ions Covalent Sharing of electrons Metallic Metal atoms bonded to several other atoms Covalent Bonding • A bond where electrons from each atom are shared

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Chapter 8 and 9 Study Guide . Chapter 8: Ionic Bonding Chapter 9: Covalent Bonding What is a Chemical Bond? The electrostatic forces holding atoms together to form a molecule (These Chemical bonds are called: Intramolecular Bonds). These bonds can be a give-take relationship between two atoms (ionic bonds) or a shared electron relationship between two atoms (covalent bonds).

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