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Unit 4 "Atomic Structure ... Dalton's Atomic Theory (experiment based!) 3) Atoms of different elements combine in simple whole-number ratios to form chemical compounds 4) In chemical reactions, atoms are combined, ... Chapter 4 Atomic Structure Author:

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model of atom, atomic mass, mass

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Atoms—and the protons, neutrons, and electrons that compose them—are extremely small. For example, a carbon atom weighs less than 2×10^{-23} g, and an electron has a charge of less

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than 2×10^{-19} C (coulomb). When describing the properties of tiny objects such as atoms, we use appropriately small units of measure, such as the atomic mass unit (amu) and the fundamental unit of ...

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Introduction to Periodic Table B. Orbits
C. Electronic configuration D. Valence
shell (Last shell of the atom) E. Valence
electrons (Electrons present in ...

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presented and connected to the ultimate “cheat sheet” in chemistry, the periodic table of elements.

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Chapter 4 Atomic Structure 33 SECTION
4.1 DEFINING THE ATOM (pages
101-103) This section describes early

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atomic theories of matter and provides ways to understand the tiny size of individual atoms. Early Models of the Atom (pages 101-102) 1. Democritus, who lived in Greece during the fourth century B.C.,

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chapter 4: the structure of the atom

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Although the ancient Greeks had many creative ideas about matter, our current understanding is based upon John Dalton's atomic theory, published in 1803. Essentially, matter is made up of atoms, which are indivisible and indestructible.

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